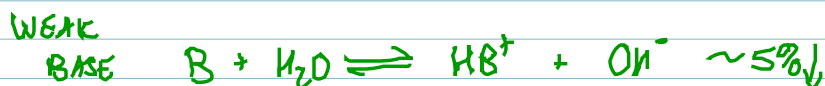
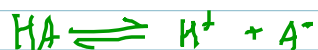


M = alkali metals, Ba



- amines $C_xH_yNH_2$ - conjugates of weak acids
- insoluble hydroxides

WEAK ACID



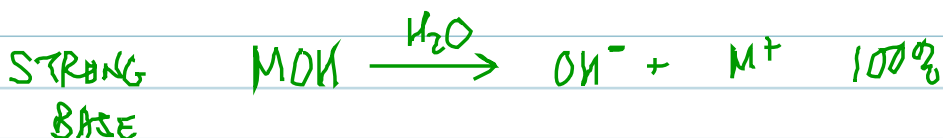
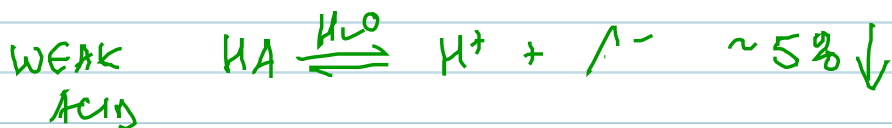
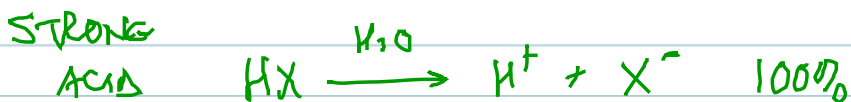
$$K_a = \frac{[H^+][A^-]}{[HA]}$$

↑ "acid ionization constant"

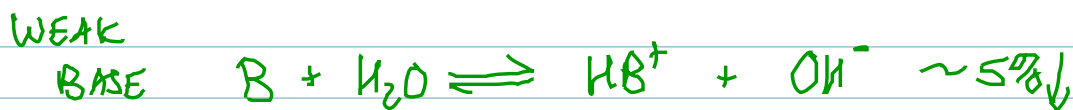
The larger the K_a , the stronger the acid

$K_a \geq 1 \Rightarrow$ STRONG ACID

$K_a \leq 10^{-3} \Rightarrow$ WEAK ACID



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